Ionic Bonding HW

Read the Chemistry flexbook pages 241-246 and answer the following questions below.

1. Draw an electron dot diagram for the following atoms and ions.

a. S

b. S2−

c. Cs

d. Cs+

e. P

f. P3−

g. Sn

h. Sn2+

2. Which of the following pairs of atoms could be expected to combine chemically to form an ionic compound? Explain.

a. Li and O

b. N and H

c. Al and S

d. Cl and F

e. Sr and Br

f. K and He

g. Li and Cl

h. Na and I

3. Use electron dot diagrams to demonstrate the formation of ionic compounds involving the following elements. Use arrows to show the transfer of electron(s) from one atom to another.

a. K and O

b. Ca and N

c. Ba and S

4. Write the formula units (compounds) for each of the ionic compounds from question number 3.

5. Write the correct chemical formula for the compounds formed from each pair of ions.

a. K+, S2-

b. Ca2+, O2-

c. Na+, O2-

d. Al3+, N3-